

Mole Review Problems

Block _____

Solve the following problems and SHOW YOUR WORK

1. How many moles of ATOMS are there in
- 1.20×10^{25}
- methane (
- CH_4
-) molecules?

$$1.20 \times 10^{25} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) \left(\frac{5 \text{ atoms}}{1 \text{ molecule}} \right) = \underline{99.7 \text{ moles of atoms}}$$

2. Calculate the molarity of a solution formed by dissolving 32.5 g of NaCl in sufficient water to make 2.5 L of solution.

$$\text{FW NaCl} = 23.0 + 35.5 = 58.5 \text{ g/mol}$$

$$32.5 \text{ g} \div 58.5 \text{ g/mol} = 0.556 \text{ mol}$$

$$M = \frac{n}{V} = \frac{0.556 \text{ mol}}{2.5 \text{ L}} = \underline{0.22 \text{ M}}$$

3. How many ATOMS are contained in 32.0 grams of gaseous oxygen?

$$\text{FW O}_2 = (16.0) \times 2 = 32.0 \text{ g/mol}$$

$$32.0 \text{ g} \div 32.0 \text{ g/mol} = 1.00 \text{ mol O}_2$$

$$1.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \times \frac{2 \text{ O atoms}}{1 \text{ O}_2 \text{ molecule}} = \underline{1.20 \times 10^{24} \text{ atoms}}$$

4. 0.500 mole of an element has a mass of 31.75 grams. Identify the element.

$$\text{FW} = \frac{31.75 \text{ g}}{0.500 \text{ mol}} = \underline{63.5 \text{ g/mol}} \quad \underline{\text{The element is Copper.}}$$

5. Calculate the number of moles of neon gas in 54 litres at STP.

$$54 \text{ L} \div 22.4 \text{ L/mol} = \underline{2.4 \text{ mol}}$$

6. A sample of carbon reacts with a 9.2 gram sample of hydrogen to form a gaseous compound that has a mass of 120.0 grams. What is the molecular formula of this hydrocarbon, which has a molar mass of 26.0 grams?

$$\text{H: } 9.2 \text{ g} \div 1.0 \text{ g/mol} = 9.2 \text{ moles H}$$

$$\text{C: } (120.0 \text{ g} - 9.2 \text{ g}) \div 12.0 \text{ g/mol} = 9.2 \text{ moles C}$$

$$\text{Mole Ratio } \text{C} : \text{H} = 1 : 1$$

$$\text{Empirical formula is } \text{CH} \rightarrow \text{molar mass } 12.0 + 1.0 = 13.0$$

$$\text{molecular mass is } 26.0 \rightarrow \underline{\text{so formula is } \text{C}_2\text{H}_2}$$

7. What is the mass of
- 2.11×10^{24}
- molecules of chlorine gas?

$$2.11 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = 3.50 \text{ moles Cl}_2$$

$$\text{FW Cl}_2 = 2(35.5) = 71.0 \text{ g/mol}$$

$$3.50 \text{ moles} \times 71.0 \text{ g/mol} = \underline{249 \text{ g}}$$

Skill Practice 51 Answers Chemistry

Michael Mosher, Paul Kelter



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MOS 51R, Interior Electrician, Skill Level 1,1988

Ebbing Darrell D. Ebbing,1993

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Chemical Dependency Counseling Robert R. Perkinson,2011-06-17 This volume is the most comprehensive guide for counselors and front line professionals who work with the chemically dependant in a variety of settings It is a basic introduction that guides the counselor through treatment from A to Z Chapters cover the gamut of treatment issues including developing the therapeutic alliance screening detoxification biopsychosocial assessment dual diagnosis patient orientation treatment planning individual therapy group therapy case management crisis intervention referral record keeping discharge planning and referral Author Robert R Perkinson provides case studies and step by step instructions with clear explanations and procedures that counselors need to use in all phases of patient care One of the key features of this volume is the inclusion of over 50 appendices which includes screening questionnaires scales history assessments personal recovery plans in patient and out patient schedules and many many more The new edition includes more appendices and updates of existing ones Inclusion of issues of diversity and working with diverse clients Discussion of new treatment for addiction with detailed plan New discussion of relapse triggers including negative emotions social pressure interpersonal conflict and the recovery community New discussion of co occurring disorders

Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998

Holt Chemistry R. Thomas Myers,2004

Common Core Science 4 Today, Grade 5,2014-05-15 Common Core Science 4 Today Daily Skill Practice provides the perfect standards based activities for each day of the week Reinforce science topics and the math and language arts Common Core State Standards all year long in only 10 minutes a day Weeks are separated by science topic so they may be completed in the order that best complements your science curriculum Review essential skills during a four day period and assess on the fifth day for easy progress monitoring Common Core Science 4 Today series for kindergarten through fifth grade covers 40 weeks of science topics with engaging cross curricular activities

Common Core Science 4 Today includes a Common Core Standards Alignment Matrix and shows the standards covered on the assessment for the week for easy planning and documentation Common Core Science 4 Today will make integrating science practice into daily classroom instruction a breeze

Drug Testing John Fay, 2013-10-22 Drug Testing focuses on the contributions of drug testing in the identification of the state of drug abuse and the implementation of policies and regulations making drug use unacceptable either socially or in the workplace The book first gives an overview of drug testing policy facilitating drug testing and treatment challenges to drug testing programs and state local and federal legislation on drug testing The text then defines the problem of drug abuse including dimensions of abuse defining the problem and the needs of employers employing a drug abuse survey and treatment modalities The manuscript ponders on the development of a drug testing policy education of employees and training of supervisors Discussions focus on the nature of learning principles of supervision training process program implementation basic assumptions final policy content and problematic issues The implementation of a drug testing program collection of specimens for drug testing and understanding the drug testing program are also underscored The publication is a valuable reference for readers interested in drug testing policies treatment and effects of drug abuse in the workplace

Cumulated Index Medicus, 1974 Objective NCERT Xtract Chemistry for NEET/ JEE Main 5th Edition Disha Experts, **Pearson Chemistry** Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, 2012

Multiple Representations in Chemical Education John K. Gilbert, David Treagust, 2009-02-28 Chemistry seeks to provide qualitative and quantitative explanations for the observed behaviour of elements and their compounds Doing so involves making use of three types of representation the macro the empirical properties of substances the sub micro the natures of the entities giving rise to those properties and the symbolic the number of entities involved in any changes that take place Although understanding this triplet relationship is a key aspect of chemical education there is considerable evidence that students find great difficulty in achieving mastery of the ideas involved In bringing together the work of leading chemistry educators who are researching the triplet relationship at the secondary and university levels the book discusses the learning involved the problems that students encounter and successful approaches to teaching Based on the reported research the editors argue for a coherent model for understanding the triplet relationship in chemical education

ACCA Approved - Strategic Business Leader (SBL) (for Sept 2018, Dec 2018, Mar 2019 & June 2019 exams) Becker Professional Education, 2017-10-01 ACCA Approved and valid for Sept 2018 Dec 2018 Mar 2019 and June 2019 exams Becker's Revision Question Bank has been approved and quality assured by ACCA's examining team and includes Professional skills practice Technical practice questions Six exam style and standard case studies including ACCA's two specimen exams Two Mock Exams Model answers include tutorial notes and professional skills marking guidance

An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or

everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Chemical Misconceptions Keith Taber,2002 Part one includes information on some of the key alternative conceptions that have been uncovered by research and general ideas for helping students with the development of scientific conceptions

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